**SAC: Experiment - Properties of ethanol**

**Preliminary discussion**

Ethanol can be produced naturally from the fermentation of sugar. There are not many countries in the world that do not have an alcohol industry. Ethanol, however, is a very useful chemical in its own right. It is a popular solvent and it is the starting point for the manufacture of other chemicals.

**Part A: Fractional distillation**

**Aim**: To obtain samples of ethanol using fractional distillation.

**Materials**

Quickfit apparatus as shown below – a measuring cylinder can be used instead of a receiving flask

Bottle of port or similar alcohol of content around 18% alcohol and red in colour

3 x 100 mL measuring cylinders

**Method**:

1**.** Measure 50 mL of wine and add it to the boiling flask

2. Add a few boiling chips

3. Have your teacher check your apparatus before heating

4. Heat gently

5. The first distillate should be obtained at a temperature around 78 0C. Catch a sample of

this in a 100 mL measuring cylinder.

8**.** Once the temperature reaches 90 0C, stop heating.



**Part B: Testing your fraction**

**Aim**: To test the properties of the fraction

**Materials**

Watchglasses

1 mL pipette

Balance

**Procedure**

**Determine the density of the alcohol fraction using the formula density = mass/volume**

**1.** Weigh a watchglass

**2.** Add 2 mL of the first alcohol fraction (measured accurately) to the watchglass and reweigh it.

**3.** Calculate the density of each sample

**Flammability**

Leave a small sample of the fraction on the watchglass

Try to set each fraction alight using a match.

Record whether the sample will burn or not.

**Solvent**

Use a large filter paper and draw any symbols on it using a series of dark biros.

Drop some of the fraction onto the symbols and observe the impact of the ethanol on the ink

**Solubility in water: Class demonstration**

Add 50 mL of water to a measuring cylinder.

Add 50 mL of ethanol (methylated spirits is a cheaper option) to another measuring cylinder.

Pour the ethanol into the water.

Observe the volume of the final solution

**Part C: Chemical reaction of ethanol**

**Aim**: To investigate the reaction of ethanol with potassium dichromate

**Materials:**

1 M sulfuric acid

0.5 M K2Cr2O7 solution

Test tube

**Procedure**

Add to a test tube – 5 drops of ethanol, 10 drops of 1M sulfuric acid and 5 drops of potassium dichromate solution.

Heat in a beaker of hot water until a colour change occurs. Record your observations.

**Questions**

Appropriate set out of results and observations 8 marks

*To include measurements and calculation of density and observations from each section*.

**Part A**

Ethanol boils at 78.5 0C and water at 100 0C. Fractional distillation should see the ethanol vaporise first and it can be collected before the water boils. Simple fractional distillation, however, does not lead to 100% separation of both components.

1. Draw a structural diagram of ethanol.

1 mark

1. Explain clearly why water has a higher boiling point than ethanol

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3 marks

1. What did you notice about the colour of the fractions compared to the original alcohol?

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2 marks

**4**. What safety precautions should be taken during distillation and why?

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3 marks

**Part B**: Testing the ethanol

**5, a**. What was the density value you obtained from your sample?

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**b**. The density of pure ethanol is listed as 0.78 g mL-1. What conclusion can you draw about the purity of your ethanol sample?

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1 + 2 = 3 marks

**6**. Write a balanced equation for the complete combustion of ethanol.

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2 marks

**7. a**. Describe the effect of dripping ethanol onto dark biro.

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**b**. What property of ethanol is this highlighting?

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2 + 1 = 3 marks

**Part C**: **Reactions of ethanol**

**8. a**. Write a balanced half equation for the reaction of ethanol to ethanoic acid.

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**b**. Write a balanced half equation for the reaction of dichromate ions to Cr3+ ions.

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**c**. Identify the oxidant and the reductant in this reaction.

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**d**. Write a balanced overall equation for the reaction.

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**e**. How do you know when all the ethanol has reacted?

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1 + 1 + 2 + 1 + 1 = 6 marks

**9. a**. What was the final volume when you added the 50 mL of water to 50 mL of water?

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**b**. Suggest reasons why the volume is not 100 mL. Use your knowledge of bonding in

your answer.

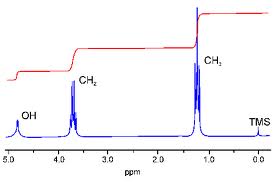
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1 + 2 = 3 marks

**Part D:**

**10.** An NMR spectrum for ethanol is shown below.

Explain each feature of this NMR

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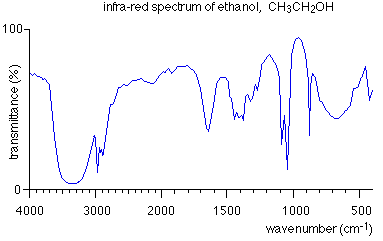
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5 marks

**11.** An infrared spectrum of ethanol is shown.



Explain what information can be ascertained from this spectrum.

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3 marks

**12**. There are several possible pathways for the production of ethanol.

**a**. **i**. Write a balanced equation for the production of ethanol from fermentation of glucose.

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**ii**. Can ethanol produced this way be referred to as bioethanol? Explain your answer.

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1 + 2 = 3 marks

**b. i**. Write a balanced equation for the production of ethanol from an alkene.

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**ii**. What category of a reaction is this? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

2 + 1 = 3 marks

**c**. Write balanced equations for the production of ethanol from an alkane in two steps.

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2 marks