**CHEMISTRY SAC**: Analysis of Data AOS2

**Aim**: To analyse a series of equilibrium data scenarios

**Case 1**: Equilibrium constant expression

Hydrogen and iodine gases can form an equilibrium with hydrogen iodide. The equation is

H2(g) + I2(g) ⇌ 2HI(g)

Four different equilibrium mixtures of these gases are formed at 25 0C and the concentration of each species is listed below.

|  |  |  |
| --- | --- | --- |
| **H2** | **I2** | **HI** |
| 0.60 | 0.40 | 1.70 |
| 0.80 | 0.50 | 2.19 |
| 1.00 | 0.50 | 2.45 |
| 1.20 | 0.70 | 3.17 |

1. Write an equilibrium expression for this mixture. 1 mark

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2. Use the data provided to prove that this data does in fact produce a constant when substituted into the

equilibrium expression. 3 marks

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3. Use the data provided to prove that the expression  does not produce a constant. 2 marks

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4. A new mixture of the above species is prepared at 25 0C and the concentration of each species is listed below:

[H2] = 5.0 M [I2] = 5.0 M [HI] = 5.0 M 3 marks

Is this mixture at equilibrium? If not, in which direction does it need to move? Explain your answer.

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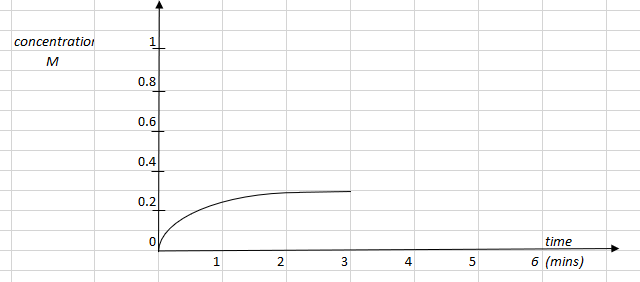
**Case 2**: Equilibrium graphs

Nitrogen oxide and oxygen form an equilibrium with nitrogen dioxide.

2NO(g) + O2(g) ⇌ 2NO2(g)

1.1 mole of NO and 1.0 mole of O2 gases are mixed in an empty 1.0 L reactor.

The graph below shows the concentration of NO2 over the first 3 minutes.



5. Draw in the concentrations of both NO and O2 that are consistent with the starting amounts and the NO2graph.

4 marks

6. Calculate the value of *K*c. 2 marks

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7. At the 3 minute mark the volume of the container is doubled.

Use the graph to show this change and to draw in the likely response of the system over the next couple of

minutes. 4 marks

**Case 3**: Equilibrium and yield

Ethane can be ‘cracked’ to form ethene and hydrogen gases. The equation for the reaction is

C2H6(g) ⇌ C2H4(g) + H2(g) ΔH = +94 kJ mol-1

An ethene manufacturer will want to maximise the yield of ethene in this reaction.

8. For each variable, explain in detail how and why it should be manipulated to lead to a maximum yield:

a. Temperature: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 3 marks

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b. Pressure: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 3 marks

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c. Catalyst: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 2 marks

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**Case 4**: Colour change

The equilibrium between nitrogen tetroxide and nitrogen dioxide is

N2O4 (g) ⇌ NO2(g)

colourless brown

9. An equilibrium mixture of the above gases is formed. The volume is then halved.

Explain carefully how the brown intensity will change after the volume change. 3 marks

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