**SAC 2: Outcome 2 Prac report on Reaction rates**

**Reaction rates with iron solutions**

**Aim**: To investigate variables that influence the rate of a chemical reaction.

**Background**

Iron (III) nitrate solution is brown in colour. When it is mixed with sodium thiosulfate, the Fe3+ is reduced to clear Fe2+.

The rate of this reaction can be determined by drawing a cross and sitting the reaction flask on top of this cross. The time can be measured until the cross becomes visible.

 *Observe cross from above*

 *Flask sits on a cross drawn on paper*

**Apparatus**

100 mL measuring cylinder

0.2 M Fe(NO3)3

0.1 M Fe(NO3)3

0.2 M Na2S2O3

0.1 M Na2S2O3

0.1 M CuSO4

0.1 M NiSO4

Stopwatch

100 mL beaker

Thermometer

Hotplate

**Part A:** Rate and temperature

Draw a cross on a piece of paper and sit a 100 mL beaker on the cross

1. Add 30 mL of 0.1 M Na2S2O3 solution to the beaker. Record the temperature

2. Add 30 mL of 0.1 M Fe(NO3)3 solution to the beaker and start timer

3. Stop the timer as soon as the cross is visible through the liquid in the flask

4. Repeat the procedure with both liquids at 30 0C

5. Repeat the procedure with both liquids at 40 0C

6. Repeat the procedure with both liquids at 50 0C

**Part B**: Rate vs Concentration

1. Repeat the procedure at room temperature but using a 0.2 M solution of Na2S2O3

2. Repeat the procedure at room temperature but using 0.2 M solutions of both liquids

3. Repeat the procedure at room temperature but using 0.05 M solutions of both liquids

**Part C**: Rate vs Catalyst

1. Repeat the procedure at room temperature using solutions of concentration 0.1 M but this

 time add 1 drop of 0.1 M CuSO4 to the Fe(NO3)3 before the two liquids are mixed.

2. Repeat the procedure at room temperature using solutions of concentration 0.1 M but this

 time add 1 drop of 0.1 M NiSO4 to the Fe(NO3)3 before the two liquids are mixed.

Appropriate recording of measurements 8 marks consisting of;

 Labelling of data 2 marks

 Use of tables 4

 Units 2

**Questions**

**1**. The reaction is between Na2S2O3 and Fe(NO3)3. It is a redox reaction.

 **a**. Write a balanced half equation for the reduction of Fe3+

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 **b**. Write a balanced half equation for the reaction of S2O32-  to S4O62-

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 **c**. Write a balanced overall equation for this reaction.

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 **d**. Explain why the solution loses colour.

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1 + 1 + 1 + 1 = 4 marks

**Part A**

2. Graph the results of time for cross to disappear against temperature.

3 marks

**3. a**. How should the rate of reaction be changing with temperature?

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 **b**. Is your graph showing this?

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2 + 1 = 3 marks

**4.** Calculate the reciprocal of each value of time taken for the cross to disappear in part A

 Draw a graph of the reciprocal against temperature.

3 marks

**5**. Explain how this graph is a better representation of the impact of temperature upon

 reaction rate.

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 2 marks

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**6**. Should this graph pass through the origin? Explain your answer.

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3 marks

**7**. Explain why the rate of reaction changes with temperature

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2 marks

**8**. List two examples you can think of where temperature is used to change the rate of a

 reaction.

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2 marks

**Part B**

**9**. How does the rate of reaction change with concentration?

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2 marks

**10**. Explain why concentration impacts the rate of reaction.

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1 mark

**11. a**. When you double the concentration of the 30 mL of Na2S2O3, does this mean the

 concentration of the Na2S2O3 is double when the reaction occurs? Explain your

 answer using a calculation

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 **b**. Explain clearly why it is more difficult to study the impact of concentration changes

 than it is to study the impact of temperature changes

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 3 + 3 = 6 marks

**12**. List two examples you can think of where concentration changes are used to affect the

 rate of a reaction.

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2 marks

**Part C**

**13. a**. What was the impact of adding the CuSO4?

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 **b**. What was the impact of adding the NiSO4?

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1 + 1 = 2 marks

**14**. a. Explain what has happened

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 **b**. Is the impact of each solution chosen the same?

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1 + 1 = 2 marks

**15**. List two examples that you can think of where a catalyst is used to affect the rate of a

 reaction.

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2 marks

**16**. Discuss safety and chemical disposal aspects of this experiment.

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3 marks